

standard electromotive force

Quantity defined by $E^\circ = -\frac{\Delta_r G^\circ}{nF} = \frac{RT}{nF} \ln K^\circ$, where $\Delta_r G^\circ$ is the standard Gibbs energy of the cell reaction in the direction in which reduction occurs at the right-hand electrode in the diagram representing the cell ('reduction at right'), K° is the standard equilibrium constant for this reaction, n its charge number, F the Faraday constant, R the gas constant and T the thermodynamic temperature.

Source:

Green Book, 2nd ed., p. 58